

NOTES – Chemistry of Life #2

a) Explain the difference between covalent and ionic bonds.

Covalent and ionic bonds are distinct types of chemical bonds that arise from the interactions of atoms based on their relative electronegativities, which is a measure of an atom's ability to attract electrons. In a covalent bond, atoms share electrons to achieve stability. This type of bonding typically occurs between atoms with similar electronegativities, usually nonmetals, where neither atom is strong enough to completely pull the electrons away from the other. If the electronegativities are equal or very close, the electrons are shared equally, resulting in a nonpolar covalent bond, as seen in molecules like O_2 . However, when there is a small difference in electronegativity between the atoms, the electrons are shared unequally, creating a polar covalent bond, such as in water (H_2O), where the oxygen atom has a higher electronegativity than the hydrogen atoms.

In contrast, ionic bonds form between atoms with a significant difference in electronegativity. Typically, this involves a metal and a nonmetal. The atom with the lower electronegativity (usually the metal) loses electrons, becoming a positively charged cation, while the atom with the higher electronegativity (usually the nonmetal) gains those electrons, becoming a negatively charged anion. This transfer of electrons results in the formation of ions that are held together by strong electrostatic forces in an ionic bond. For example, in sodium chloride ($NaCl$), sodium (Na) has a much lower electronegativity than chlorine (Cl), leading to the complete transfer of an electron from sodium to chlorine, forming a strong ionic bond between the Na^+ and Cl^- ions. This difference in electronegativity is the key factor that distinguishes ionic bonds from covalent bonds.

b) Are hydrogen bonds true bonds?

Hydrogen bonds are not considered true chemical bonds like covalent or ionic bonds; instead, they are classified as a type of intermolecular force. A true bond involves the sharing or transfer of electrons between atoms, leading to the formation of a stable chemical structure. In contrast, a hydrogen bond is an attractive interaction that occurs between a hydrogen atom, which is covalently bonded to a highly electronegative atom (such as oxygen, nitrogen, or fluorine), and another electronegative atom in a nearby molecule or within the same molecule. This interaction is primarily electrostatic, resulting from the partial positive charge on the hydrogen atom and the partial negative charge on the electronegative atom.

Despite not being true bonds, hydrogen bonds play a crucial role in the structure and properties of many substances. For example, hydrogen bonds are responsible for the unique properties of water, such as its high boiling point, surface tension, and ability to dissolve many substances. They also contribute to the three-dimensional structures of proteins and nucleic acids like DNA, where hydrogen bonds between base pairs help stabilize the double helix. While weaker than true covalent or ionic bonds, hydrogen bonds are significant in biological systems and various chemical processes due to their collective strength when present in large numbers.

c) Explain Metallic bonds and the properties they exhibit.

Metallic bonds are a type of chemical bond found in metals, characterized by a "sea of electrons" that are free to move throughout the entire structure. In metallic bonding, atoms in a metal release some of their electrons, which become delocalized and flow freely between the positively charged metal ions. This creates a lattice of metal cations immersed in a pool of free-moving electrons. The electrostatic attraction between the positively charged ions and the negatively charged electrons holds the metal together, giving rise to the metallic bond. This unique bonding allows metals to exhibit several distinct properties.

One of the key properties of metallic bonds is electrical conductivity. The free electrons in the metallic structure can move easily under an electric field, allowing metals to conduct electricity efficiently. Additionally, metallic bonds contribute to the malleability and ductility of metals, as the layers of atoms in a metal can slide past one another without breaking the bond, enabling the metal to be shaped or drawn into wires. The strength of the metallic bond also accounts for the high melting and boiling points of most metals, as a significant amount of energy is required to break these bonds. Furthermore, the luster and reflectivity of metals arise from the interaction of light with the free electrons on the surface, which can absorb and re-emit photons, giving metals their shiny appearance.

d) What is a polar molecule, and why is water considered polar?

A polar molecule is one in which there is an unequal distribution of electron density, leading to the formation of regions with partial positive and partial negative charges within the molecule. This occurs due to differences in electronegativity between the atoms involved in the molecule. Electronegativity is a measure of how strongly an atom attracts electrons in a bond. When atoms with different electronegativities form a covalent bond, the shared electrons tend to be closer to the more electronegative atom, creating a charge imbalance. As a result, one end of the molecule becomes slightly negative, while the other end becomes slightly positive, giving rise to a dipole moment. Polar molecules are characterized by this separation of charge and often exhibit interactions such as hydrogen bonding and dipole-dipole interactions, which play a crucial role in their physical and chemical properties.

Water (H₂O) is a classic example of a polar molecule, primarily due to the significant difference in electronegativity between the oxygen and hydrogen atoms and the molecule's bent shape. Oxygen is more electronegative than hydrogen, meaning it has a stronger tendency to attract the shared electrons in the O-H bonds. As a result, the electrons spend more time closer to the oxygen atom, giving it a partial negative charge (δ^-). Conversely, the hydrogen atoms, with less access to these shared electrons, acquire a partial positive charge (δ^+). The shape of the water molecule, which is bent rather than linear, ensures that these charges are asymmetrically distributed, reinforcing the polarity of the molecule.

The concept of dipole moment is central to understanding molecular polarity. The dipole moment is a vector quantity that represents the magnitude and direction of the separation of charge within a molecule. It is calculated as the product of the charge difference and the distance between the charges. In mathematical terms, the dipole moment (μ) is given by $\mu = q * d$, where q is the magnitude of the charge and d is the distance between the centers of positive and negative charge. A larger dipole moment indicates a greater separation of charge and a stronger polarity. The direction of the dipole moment points from the positive to the negative charge, which, in the case of water, points from the hydrogen atoms towards the oxygen atom.

Water's dipole moment is significant due to both the large electronegativity difference between oxygen and hydrogen and the bent shape of the molecule. This strong dipole moment makes water a highly polar molecule, which accounts for many of its unique properties. For example, water's polarity allows it to dissolve many ionic and polar substances, making it an excellent solvent. The polarity also enables water molecules to form hydrogen bonds with each other, contributing to water's high boiling point, surface tension, and ability to support life through processes such as capillary action and temperature regulation in living organisms. The concept of the dipole moment, therefore, not only explains why water is polar but also helps to understand the broader implications of this polarity in chemistry and biology.

5) What are the major types of carbon-based macromolecules found in living organisms?

Living organisms are primarily composed of four major types of carbon-based macromolecules: carbohydrates, lipids, proteins, and nucleic acids. These macromolecules are crucial for various biological functions and form the structural and functional backbone of cells.

Carbohydrates are composed of carbon, hydrogen, and oxygen, typically in a ratio of 1:2:1. They serve as the primary energy source for most organisms. Simple carbohydrates like glucose provide immediate energy, while complex carbohydrates like starch and glycogen are used for long-term energy storage. Carbohydrates also play a role in cell structure, particularly in the cell walls of plants (cellulose) and the exoskeletons of arthropods (chitin).

Lipids are a diverse group of hydrophobic molecules, including fats, oils, phospholipids, and steroids. They are primarily composed of carbon and hydrogen, with a small amount of oxygen. Lipids serve as long-term energy storage molecules, provide insulation and protection for organs, and are key components of cellular membranes. Phospholipids, in particular, form the bilayer structure of cell membranes, contributing to their fluidity and integrity.

Proteins are complex molecules made up of amino acids linked by peptide bonds. They serve a wide range of functions, including catalyzing biochemical reactions (enzymes), providing structural support (collagen), and facilitating communication between cells (hormones and receptors). The specific sequence and structure of a protein determine its function, which is vital for maintaining cellular processes.

Nucleic acids, including DNA and RNA, are polymers of nucleotides that store and transmit genetic information. DNA contains the instructions for building proteins, while RNA is involved in translating these instructions during protein synthesis. Nucleic acids are essential for heredity, gene expression, and the regulation of cellular activities.

6) Describe the process of dehydration synthesis and its role in forming macromolecules.

Dehydration synthesis, also known as condensation reaction, is a chemical process by which two molecules are joined together with the removal of a water molecule. This process is fundamental in the formation of macromolecules such as carbohydrates, proteins, and nucleic acids.

During dehydration synthesis, a hydroxyl group (—OH) from one molecule and a hydrogen atom (—H) from another are removed, forming water (H_2O). The remaining parts of the molecules are covalently bonded, resulting in the formation of a larger molecule. For example, in the formation of a disaccharide like sucrose, a glucose and a fructose molecule are linked together by removing a water molecule.

In proteins, dehydration synthesis occurs during the formation of peptide bonds between amino acids. Each amino acid has an amino group (—NH_2) and a carboxyl group (—COOH). When these groups react, a water molecule is released, and a peptide bond is formed, linking the amino acids into a polypeptide chain.

Similarly, in nucleic acids, dehydration synthesis is responsible for forming phosphodiester bonds between nucleotides, which link the sugar of one nucleotide to the phosphate group of the next. This process is critical for the formation of the DNA and RNA backbones, which are essential for storing and transmitting genetic information.

7) How do enzymes lower the activation energy of a reaction?

Enzymes are biological catalysts that speed up chemical reactions by lowering the activation energy required for the reaction to proceed. Activation energy is the energy barrier that must be overcome for reactants to be converted into products. By lowering this barrier, enzymes enable reactions to occur more rapidly and efficiently under physiological conditions.

Enzymes lower activation energy through several mechanisms. One key method is by providing an optimal environment for the reaction. The enzyme's active site, where the substrate binds, creates a microenvironment that can favor the reaction by aligning reactive groups, stabilizing transition states, or providing acidic or basic conditions conducive to the reaction.

Another way enzymes lower activation energy is by inducing strain in the substrate. When the substrate binds to the active site, the enzyme may change shape slightly, a process known as induced fit. This conformational change can stress specific bonds within the substrate, making them easier to break or rearrange, thus lowering the energy needed for the reaction.

Additionally, enzymes can participate directly in the reaction by temporarily forming covalent bonds with the substrate or by donating and accepting electrons, which helps facilitate the transformation of the substrate into the product. These interactions between the enzyme and substrate stabilize the transition state and reduce the activation energy, making the reaction proceed more quickly.

8) Explain the process of protein folding and the factors that influence it.

Protein folding is the process by which a polypeptide chain acquires its functional three-dimensional structure. This process is critical for the protein's biological activity, as the specific shape of a protein determines its function. Protein folding is influenced by various factors, including the sequence of amino acids, the cellular environment, and molecular chaperones.

The primary structure of a protein, which is the linear sequence of amino acids, dictates how the protein will fold. The sequence determines the distribution of hydrophobic and hydrophilic residues, which play a crucial role in the folding process. Hydrophobic amino acids tend to cluster in the interior of the protein, away from the aqueous environment, while hydrophilic amino acids are more likely to be exposed on the surface.

During folding, the polypeptide chain forms secondary structures, such as alpha helices and beta sheets, stabilized by hydrogen bonds. These secondary structures further fold into a tertiary structure, which is the overall three-dimensional shape of the protein. In some cases, multiple polypeptide chains (subunits) come together to form a quaternary structure.

Several factors can influence protein folding, including the presence of molecular chaperones, which assist in the correct folding of proteins by preventing aggregation and guiding the protein to its native state. The cellular environment, such as pH, temperature, and ionic strength, also affects folding. Misfolding can lead to non-functional proteins or aggregates, which can be harmful to the cell.